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Chapter 13 Test Chemical Reactions

Chapter 13 Chemical Reactions-Chemical Reactions. STUDY. PLAY. enzyme. an example of a catalyst such as the one in stomach acid or an apple. catalyst. a substance that speeds up a chemical reaction, but is not consumed. chemical reaction. changes the way in which atoms are arranged by breaking and forming bonds to form different compounds.

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Chapter 13 Chemical Reactions 1. Reactants disappear or are diminished 2. New substances are formed as products. These products have different chemical/physical properties from the original... 3. Energy is either released or absorbed and will be accompanied by heat, light, electricity, or sound.

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Chapter 13 chemical reactions. Chemical reaction. Precipitate. Chemical formula. Chemical equation. The process by which one or more substances change to produce.... A solid that forms from a solution during a chemical reaction. A combination of chemical symbols and numbers to represent a s....

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346 CHAPTER 13: CHEMICAL KINETICS 13.27We know that half of the substance decomposes in a time equal to the half-life, t1/2. This leaves half of the compound. Half of what is left decomposes in a time equal to another half-life, so that only one quarter of the original compound remains.

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13. It takes 42.0 min for the concentration of a reactant in a first-order reaction to drop from 0.45 M to 0.32 M at 25°C. How long will it take for the reaction to be 90% complete? A. 13.0 min B. 86.0 min C. 137 min D. 222 min E. 284 min

Chapter 13 Chemical Kinetics

The whole series of reactions involves many chemicals, enzymes, breaking and making chemical bonds, the transfer of electrons and H + ions, and other chemical processes. The elucidation of the actual steps of photosynthesis—a process still unduplicated artificially—is a major achievement of modern chemistry.

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Lets apply the segregation model to an ideal PFR and see if we get the same result for conversion as we did in Chapter 4. Solve for X(t) for a first order reaction in a batch reactor. For the batch reactor the conversion-time relationship is Calculate the mean conversion which is the same conversion one finds from a mole balance (Chapter 4)

Chapter 13 Summary Notes

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